

Nitration of Some 2,3-Dihydro-1,4-Diazepinium Perchlorates

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NUCLEOPHILIC, free-radical, and electrophilic attack on the 2,3-dihydro-1,4-diazepine nucleus by neutral and negatively charged species is known.¹ Attack by a positively charged species however, has not been reported. It has been clearly demonstrated² that the nitrating species in nitric acid-sulphuric acid mixtures is NO_2^+ . I now report substitution of the 2,3-dihydro-1,4-diazepine nucleus by positively charged species.

2,3-Dihydro-5,7-dimethyl-1,4-diazepinium perchlorate (I) and 2,3-dihydro-1,4,5,7-tetramethyl-1,4-diazepinium perchlorate (II) undergo mononitration in nitric acid-sulphuric acid. The perchlorates (I) and (II) were heated at 70° for 2 hr. with nitric acid in 90% sulphuric acid. Yellow perchlorates (III) and (IV) were obtained on pouring the reaction mixture on to crushed ice, and had melting points (needles from water) 240–242° (III) and 231–232° (IV).

(II) has λ_{max} 318 m μ and ϵ_{max} 13,800 \pm 50. (IV) has λ_{max} 328 m μ and ϵ_{max} 18,000 \pm 50. The infrared spectra of (III) and (IV) are very similar to those of the parent compounds, but show intense absorptions at 1330 cm^{-1} which is generally associated with a nitro-group. The ¹H n.m.r. spectra of (III) and (IV) consist of three singlets, one less than (I) and (II) (Table).

The protons at τ 2.25 (I) and τ 2.8 (III) underwent deuterium exchange rapidly whilst those at τ 4.9 (I) and (II) underwent slow exchange [$k = 1.02 \pm 0.05 \times 10^{-6} \text{ sec}^{-1}$ for the pseudo-first-order exchange of the C-6 proton in (I)].³

TABLE

Proton	Solvent	Chemical shift τ			
		(I)	(II)	(III)	(IV)
N-H	MeCN	2.25	—	2.8	—
N-CH ₃	Me ₂ SO	—	6.65	—	6.65
H-6	Me ₂ SO	4.9	4.9	—	—
2(3)-CH ₂	Me ₂ SO	6.45	6.45	6.40	6.40
5(7)-CH ₃	Me ₂ SO	7.7	7.7	7.55	7.50

From a comparison of the spectral data and analytical figures it is concluded that (III) and (IV) are respectively 2,3-dihydro-5,7-dimethyl-6-nitro-1,4-diazepinium perchlorate and 2,3-dihydro-6-nitro-1,4,5,7-tetramethyl-1,4-diazepinium perchlorate.

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² E. D. Hughes and C. K. Ingold, with R. J. Gillespie, D. J. Millen *et al.*, *J. Chem. Soc.*, 1950, 2400–2684.

³ C. Barnett and J. Warkentin, Paper in preparation.